University Chemistry II CHEM1123 (Lecture)

Spring 2024

Class Time and Location: **MW** 12:30 - 1:50, **Room**: S235
Instructor: Gaumani Gyanwali – Ph. D.: OSU-Stillwater; MS, BS: Tribhuvan University
Office Location: S244
Office Hours: Posted on the office door
Contact Information: Office: (479)-394-7622, Ext. 1339
Cell: (479)-234-8727 **Email**: gyanwali@uarichmountain.edu

Prerequisite: Successful completion of 1and a half units of high school algebra, OR a co-requisite MTH203 or Permission of Instructor

Resources:

Textbook: Chemistry Zumdahl.and Zumdahl: Brooks Cole Cengage, 2010. *ISBN 978-0-547-12532-9*

Required Materials:

1. Calculator capable of roots, exponents, and logarithms.
2. One 1-Inch binder for class worksheets; Miscellaneous: pens, notes, and pencils.

Course Description: This course is a 3 credit hours course and is designed to provide the basic chemistry background necessary for science majors such as Physics, Chemistry, Biology, Biochemistry, Pharmacy, Forestry, Medicines, Engineering, etc. The theories of chemistry based on fundamental principles include colligative properties, chemical kinetics, chemical equilibrium, acid-base theory and equilibrium, solubility and complex ion equilibria, thermodynamics, and redox reactions and electrochemistry. Must take the laboratory course CHEM1121 concurrently with this course to have full 4 credit hours.

Program Outcomes: Students will increase their
-writing, reading, speaking, and listening skills,
-competency in mathematics, utilizing their analytical and occupational skills to solve real life problems,
-understanding of different cultures and societies across the globe and how both have changed over time,
-understanding of basic principles and their application in the field of physical sciences and biological sciences, and
-technological and research skills.

Rationale: Through the study of chemistry, problem solving experiences, and laboratory activity, the student develops an understanding of matter and the laws that explain the interactions between matters, including the energy changes involved. The gained knowledge is applicable in upper level chemistry courses.

Course Objectives: Upon completion of University Chemistry II, the students will increase their understanding of solubility and solution composition using various controlling factors affecting them, different colligative properties and their dependence on the solute present,

kinetics of a reaction by applying rate laws, perform calculations based on the information from the rate laws, equilibrium constant and the equilibrium expression and their application to solve problems quantitatively and qualitatively, pH of strong and weak acids and bases, acidic and basic properties of salts and oxides, titrations between acids and base involving pH measurements applying common ion effect, difference between solubility and solubility product constant, relationship among three fundamental quantities of thermodynamics, electrochemical reactions as redox reactions, fuel cell half reactions.

Topical/Unit Outline:

I. Properties of Solutions:
   a. Solution composition
   b. Energies of solution formation
   c. Factors affecting solubility
   d. Vapor pressure of solutions
   e. Boiling point elevation and freezing point depression
   f. Osmotic pressure
   g. Colligative properties of electrolyte solutions
   h. Colloids

II. Chemical Kinetics
   a. Reaction rates
   b. Rate laws
   c. Integrated rate laws
   d. Reaction mechanisms
   e. Activation energy and transition state
   f. Catalysis

III. Chemical Equilibrium
   a. Equilibrium condition
   b. Equilibrium constant
   c. Equilibrium expression involving pressure
   d. Heterogeneous equilibrium
   e. Application of equilibrium constant
   f. Solving equilibrium problems
   g. Le Chatelier's Principle

IV. Acid-Base Theory
   a. Nature of acids and bases
   b. Strong and weak acids
   c. The pH-scale
   d. Calculation of pH of strong and weak acids
   e. Calculation of pH of strong and weak bases
   f. Polyprotic acids
   g. Acidic and basic properties of salts
   h. Acidic and basic oxides
   i. Lewis acid-base model

V. Acid-Base Equilibrium
   a. Common ion effect
   b. Buffer solutions
   c. Buffering capacity
   d. Titrations and pH curves
   e. Acid base indicators

VI. Solubility and complex -ion equilibria
   a. Solubility product and solubility product constant

b. Calculation of solubility from K_{sp}

VII. Thermodynamics
   a. Spontaneous process and entropy
   b. Entropy and 2nd Law of Thermodynamics
   c. Effect of temperature
   d. Free energy
   e. Entropy change in a chemical reaction
   f. Free energy and chemical reaction
   g. Free energy and pressure
   h. Free energy and equilibrium
   i. Free energy and work

VIII. Oxidation Reduction and Electrochemistry
   a. Balancing oxidation reduction equations
   b. Galvanic cells
   c. Standard reduction potential
   d. Cell potential, electrical work, and free energy
   e. Batteries
   f. Fuel cells
   g. Electrolysis and electroplating

Tentative Schedule for Lecture: University Chemistry II

Week	**Topics/Activities by Book Chapter**	**Date/Day**
1	Syllabus, Properties of solution	1/17 W
2	*Instructor out of office (no class)* Properties of solution	*1/22 M* 1/24 W
3	Properties of solution/Chemical Kinetics	1/29 M, 1/31 W
4	Chemical Kinetics	2/5 M, 2/7 W
5	Review **Test 1**	2/12 M **2/14 W**
6	*Winter break – no class* Chemical Equilibrium	*2/19 M* 2/21 W
7	Chemical Equilibrium Acid-Base Theory	2/26 M 2/28 W
8	Acid-Base Theory Review	3/4 M 3/6 W
9	**Test 2** Acid Base Equilibrium	**3/11 M** 3/13 W
10	*Spring Break - no class*	*3/18 - 3/22*
11	Acid Base Equilibrium Solubility and Complex Ion Equilibria	3/25 M 3/27 W
12	Solubility and Complex Ion Equilibria Review	4/1 M 4/3 W
13	**Test 3** Thermodynamics	**4/8 M** 4/10 W
14	Thermodynamics /Electrochemistry	4/15 M, 4/17 W
15	Electrochemistry	4/22 M,4/24 W
16	Review **Test 4**	4/29 M **5/1 W**
17	**Final Exam**	**5/8 W**

Course Conduct: The class will be taught using a combination of lecture using *Powerpoint* presentation, solving in-class questions, demonstrations, group interaction and discussions. Students are strongly encouraged to discuss concepts and examples as opportunities are presented during instructor presentation. It is the student's responsibility to inform the instructor when a concept is confusing or otherwise unclear. Students must answer all the in-class questions before leaving the classroom and discuss with the instructor about any confusing concept right away.

You are encouraged to solve problems forming a study group and actively taking part in discussion. You may seek assistance from Student Support Services (SSS) check the concepts in homework assignments, quizzes, and exam reviews but the instructor is the primary resource in this course as tutors may not be available. You must inform me no later than one week after the classes started if you feel that the difficulty level is too high. You cannot pass the course without completing the homeworks. Expect a minimum of 12 hours of work per week for this class.

UA Rich Mountain Credit Hour Calculator: For every course credit hour, you should expect to spend at least three clock hours per week of concentrated attention on course-related work, including but not limited to time attending class, as well as out-of-class time spent completing homework, seeking help from instructor or tutors, preparing for upcoming quizzes/tests/exam, reviewing material, problem solving, and organizing your notebook. For a 3 credit-hour course, a typical student should expect to spend at least twelve hours per week dedicated to this course. This includes the time spent in class. Please see credit hour calculator attached to end of syllabus.

Course Requirements/Evaluation (Grading):

Your grade for this course comes from 4 tests (300 points, 100 points each, 1 lowest test dropped), one comprehensive final exam (100 points), 8 assignments (400 points, 50 points each), work shown on 8 assignments (50 points, 6.25 points each), and from class participation and attendance (50 points).

Four tests @100 each	= 300 (lowest one dropped after the 4th test)
One Final Exam @ 100	= 100
Attendance and Participation	= 50
8 Assignment work @ 6.25 each	= 50
8 Assignments @ 50 each	= 400
Total points possible	**= 900**

The final grade for the course will be determined by dividing the student's total number of points earned by the number of possible points. Grades are assigned on the basis of the following percentage scale, after rounding the calculated percentage to the nearest whole number:

90-100	**A**
80- 89	**B**
70- 79	**C**
60- 69	**D**
0- 59	**F**

Class Procedures/Student Requirements:

On-Campus Resources:

1. Your Blackboard will be populated with the lecture slides for each upcoming chapter. If you wish, you can print the lecture slides of the chapter.
2. You will have in-class practice worksheets to see how the instructor performs the calculations and chemical formulas. Retain those worksheets for future information and knowledge.
3. Student Support Services will have some basic information available about time management, difficulty of the course etc.

tutor.com provides 5 hours of free tutoring online for each UA Rich Mountain student per semester. Johnson Learning Commons has contract with this online site where personal tutors are available 24/7. Create an account and get help personally from home or anywhere if internet access is available. Follow the link to log in www.tutor.com/uarichmountain

Attendance Policy: Absence in this course most often results into failing grades unless covered by immediate action and make-up work. Absence is strictly prohibited unless any unavoidable circumstances. If notified of the absence, the instructor can send you the work and you can catch up with completed work. It is your full responsibility to complete the missed work and cover the missed chapter.

Assignment Due Dates: Students are expected to submit the assignments and other works on or before the assigned due dates. You will be penalized for the late works depending on the number of days and full point and importance of the works assigned.

Make-up Work Policy**:** Make up work should be competed promptly after the absence. A maximum of one week will be the guideline. Make-up exams are given only in instances of the most unavoidable, serious circumstances. Prior notice of absence from an exam is required. Upon return after an absence the student is expected to be prepared for the days' class activities.

Cell Phone Policy**:** Do not use the cellphone unless allowed by the instructor.

Extra Credit Opportunity: Optimum student involvement in the course as well as course-related activities is highly recommended and rewarded with extra credit. The value of extra credit points varies with the nature of the activity and the application of that activity in the real life as well as relation of the activity with the course. Points that can be received are unlimited. Simple idea is: the more you are involved in real-life science experiences, the more you understand and enjoy it, and you get good grades!!! Some examples of activities (but not limited to) that earn extra credit points include:

a) Complete and submit the end of chapter Exercises for each chapter.
b) Being an active member of the Science Club coordinated by the instructor
c) Organize activities within the Science Club or participate actively in the activities organized by the Club, present a project in the Science Fair in the Spring semester
d) Taking leadership in group study sessions and manage study sessions with peers
e) Organize or be involved in science and technology awareness activities and helping society to improve utilizing science knowledge
f) Actively participating in science-related meetings and conferences on/off campus

g) Any other science related activities and duties beyond the demand of the course deemed worthy of extra credit by the instructor

Academic Dishonesty: All students must do the work on worksheets, homework, exams, lab activity, and lab reports all by themselves. Cheating, copying other's answers, using other's work and changing content as well as using AI tools to create a work all are considered plagiarism and are penalized from giving a zero grade on that assignment up to a failing grade on that course, depending on the severity/repetition of the offense.

Artificial Intelligence (AI) Policy: Use of AI tools such as, Grammarly, ChatGPT, Bard, etc. in this class is allowed to verify the knowledge, look up for known constants and values, enhance the understanding of certain concept, and improve and correct writing styles and format. However, use of AI tools is prohibited during exams (for any purpose) and for creating journals and lab reports by asking those tools to develop the whole document. Prohibited activities are penalized according to Academic Dishonesty policy stated above.

Disability Policy: It is the policy of UA Rich Mountain to accommodate students with disabilities, pursuant to state and federal law. Any student with a disability who needs accommodations should contact the Associate Vice-Chancellor for Student Affairs Office, 394-7622, ext. 1460.

Arkansas Clean Air Act: In compliance with the Arkansas Clean Air Act, UA Rich Mountain is a smoke and tobacco free facility. Students and guests are to refrain from the use of all tobacco products while on campus.

Student Misconduct: Students committing misconduct are subject to disciplinary sanction(s) and/or restrictions. Misconduct or prohibited behavior includes, but is not limited to endangerment, harassment, disruption, sexual misconduct, illegal drugs and/or alcohol, fire or safety issues, gambling, property violations, computer violations or other violations as listed in the UA Rich Mountain catalog.

UA Rich Mountain Mission: UA Rich Mountain provides transformative education to all learners.

Credit Hour Requirement Calculator

Course Information				
Department, Course Number	CHEM1123(Lecture) and CHEM1121(Lab)			
Course Title	University Chemistry II			
Credit Hours	3 + 1 = 4			
Mode of Delivery	Traditional			
Weeks in Part of Term	15			
Level of Reading Material	Advanced/Scientific/Technical (2)			2
Reading Level of Students	Average Student (2)			2
Average Pages per Hour	10			
Contact Hours		**Hours/Week**	**Hours/Term**	**Total Hours**
Lecture/Recitation			45	45
Lab Contact Hours			24	24
	Pages Asg'd	**Hours/Week**	**Hours/Term**	
Syllabus Items				
Reading	380			38
Assignment		1		15
Quiz		1		15
Researching Lab questions		2		30
Reviewing for exams			10	10
				0
				0
				0
				0
				0
Total Hours				177
Required Hours				150
Difference				Ok + 27